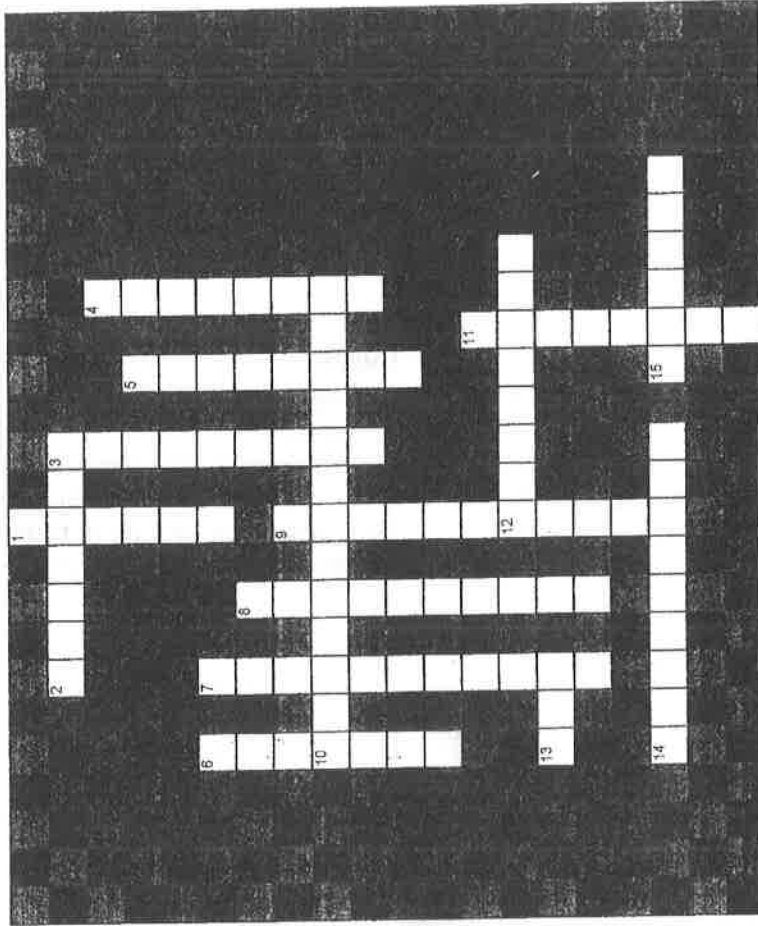


Chemistry: Matter and Change

Name: _____ Date: _____

Instructions: Complete the crossword puzzle. Use the clues to help you solve the puzzle.



Across

- _____ solution - A solution in which the solvent is water.
- _____ reaction - A chemical reaction that occurs when a single compound breaks down into two or more elements or new compounds.
- _____ ionic equation - An ionic equation that shows all the particles in a solution as they realistically exist.
- _____ ionic equation - An ionic equation that includes only the particles that participate in the reaction.

Chemistry: Matter and Change

- _____ ion - An ion that does not participate in a reaction and usually is not shown in an ionic equation.
- _____ -replacement reaction - A chemical reaction that occurs when the atoms of one element replace the atoms of another element in a compound.

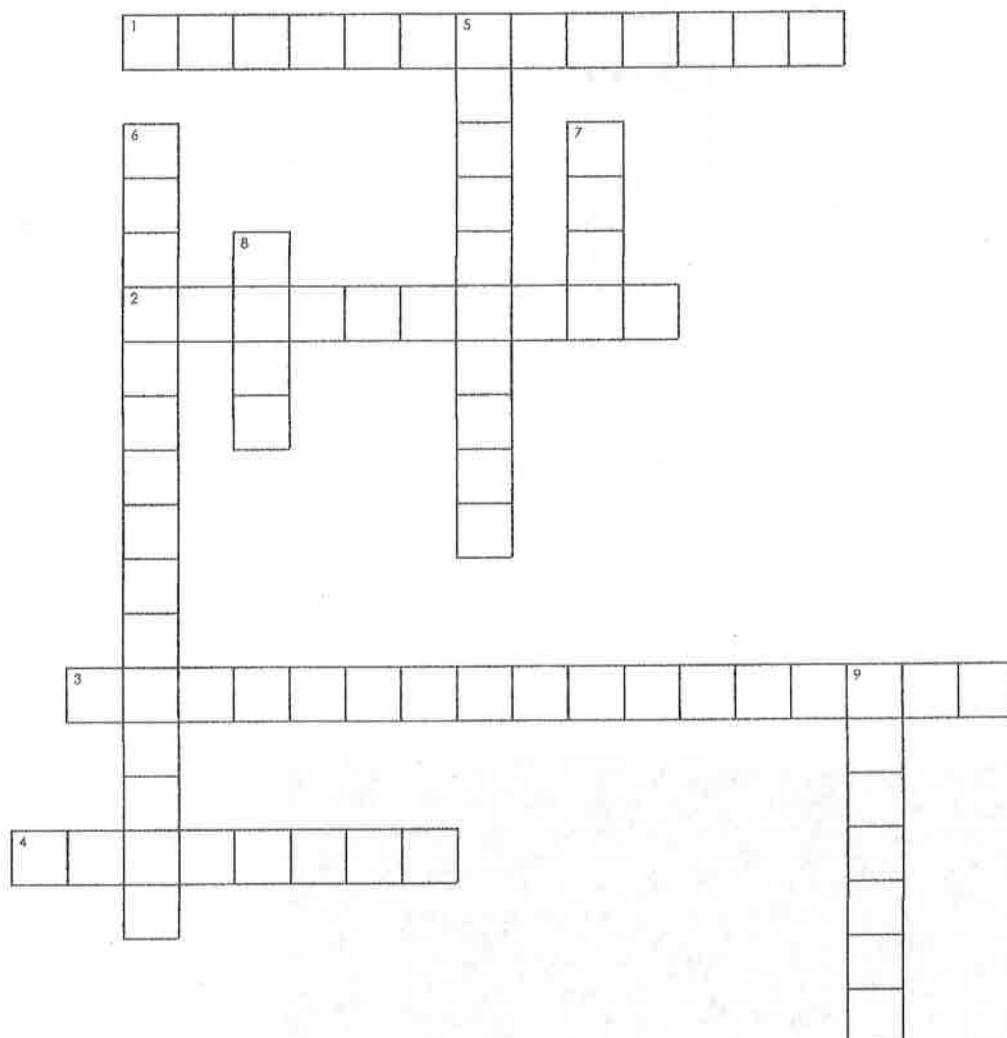
Down

- _____ -replacement reaction - A chemical reaction that involves the exchange of positive ions between two compounds dissolved in water and produces either a precipitate, a gas, or water.
- _____ reaction - A chemical reaction in which two or more substances react to yield a single product.
- The starting substance in a chemical reaction.
- Chemical _____ - A statement using chemical formulas to describe the identities and relative amounts of the reactants and products involved in the chemical reaction.
- A substance formed during a chemical reaction.
- A solid produced during a chemical reaction in a solution.
- _____ reaction - A chemical reaction that occurs when a substance reacts with oxygen, releasing energy in the form of heat and light.
- In a chemical equation, the number written in front of the formula; tells the smallest number of particles of the substance involved in the reaction.
- Chemical _____ - The process by which the atoms of one or more substances are rearranged to form different substances; indicated by changes in temperature, color, odor, and physical state.

Name: _____

Period: _____

Net Ionic Equations



Across: →

1. Precipitate that forms from BaI_2 and Cu_2SO_4
2. Result of $\text{Ni}(\text{NO}_3)_2$ and KBr
3. Precipitate that forms from K_2CO_3 and CuSO_4
4. Gas that forms when Zn reacts with HCl

Down: ↓

5. Nonmetal Spectator Ion from K_2CO_3 and CuSO_4
6. Precipitate that forms from $\text{NaOH} + \text{FeCl}_2$
7. Formula for Copper (II) Sulfate
8. Formula for Strontium Bromide
9. Dissolved in Water (in solution)