

Solubility Rules Worksheet

1. Classify each of the substances as being soluble or insoluble in water.

- a. KBr =
- b. PbCO_3 =
- c. BSO_3 =
- d. zinc hydroxide =
- e. sodium acetate =
- f. silver iodide =
- g. cadmium (II) sulfide =
- h. zinc carbonate =

- i. silver acetate =
- j. copper (II) sulfide =
- k. $\text{Mg}_3(\text{PO}_4)_2$ =
- l. KOH =
- m. NiCl_2 =
- n. NH_4OH =
- o. Hg_2SO_4 =
- p. PbI_2 =

2. Identify the two new compounds which form if the solutions, as suggested by the following table, were mixed. CIRCLE the names of the compounds which would precipitate from the solutions.

	KBr	Na_2CO_3	CaS	NH_4OH
AgNO_3				
BaCl_2				
$\text{Al}(\text{NO}_3)_3$				
CuSO_4				

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Solubility Rules Worksheet

9-20-2013

1. Name or give the chemical formula for each of the following compounds.
2. State whether they are soluble (will dissolve) or insoluble (will not dissolve) in solution. Use solubility rules.

Chemical Formula	Name	Solubility
1. $\text{NH}_4\text{C}_2\text{H}_3\text{CO}_2$		
2. $\text{Ba}(\text{OH})_2$		
3.	Iron (II) Carbonate	
4. NaOH		
5. RbNO_3		
6.	Cesium Sulfate	
7. MgSO_4		
8. ZnCl_2		
9.	Zinc Hydroxide	
10. $\text{Zn}_3(\text{PO}_4)_2$		
11. AgBr		
12. KNO_3		
13. Al_2S_3		
14.	Silver Acetate	
15. Sr_2CrO_4		
16.	Aluminum Phosphate	
17. BaSO_4		
18. $\text{Ca}(\text{OH})_2$		
19. BaCO_3		
20. MgCrO_4		
21.	Iron (III) sulfide	
22. NH_4CN		
23.	Silver Iodide	
24. Hg_2SO_4		
25.	Lithium Chloride	

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4. NaOH		
5. RbNO_3		
6.	Cesium Sulfate	
7. MgSO_4		
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