

Episode 11 - The Mole

1. Why is it important to use the correct amount of materials in a chemical reaction?
2. What names are given to the materials at the beginning and end of a chemical reaction?
3. Atoms and molecules are extremely small. How do chemists "count" them? Can you think of an everyday application of this?
4. a. What did early chemists discover about reactions involving the combination of gases?

b. How did Avogadro explain this?
5. How may a chemical equation such as $\text{H}_2 + \text{Cl}_2 \longrightarrow 2 \text{HCl}$ be interpreted?
6. What is true about the mass of a compound?
7. What is the numerical value for Avogadro's Number?
8. When the I V solutions were prepared, quality control was involved. What is quality control?
9. Why did using twice as much magnesium not produce twice as much hydrogen in the demonstration?
10. What ratio of starting materials was found to produce the best epoxy resin?

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